

Answer to Chemistry Test 2

English Name (1 each)	Symbol	State at RT Solid = s Liquid = l Gas = g (½ each)	Nature Metal = m Non-metal = n Noble gas = g Semi-metal = s (½ each)
Boron	B	s	s
Neon	Ne	g	g
Iron	Fe	s	m
Bromine	Br	l	n
Phosphorus	P	s	n
Platinum	Pt	s	m
Beryllium	Be	s	m
Fluorine	F	g	n
Barium	Ba	s	m
Argon	Ar	g	g
Iodine	I	s	n
Silver	Ag	s	m
Magnesium	Mg	s	m
Cobalt	Co	s	m
Hydrogen	H	g	n
Silicon	Si	s	s
Helium	He	g	g

English Name	Symbol (1 each)	State at RT Solid = s Liquid = l Gas = g (½ each)	Nature Metal = m Non-metal = n Noble gas = g Semi-metal = s (½ each)
Chlorine	Cl	g	n
Zinc	Zn	s	m
Carbon	C	s	n
Lithium	Li	s	m
Tin	Sn	s	m
Lead	Pb	s	m
Nitrogen	N	g	n
Potassium	K	s	m
Mercury	Hg	l	m
Nickel	Ni	s	m
Oxygen	O	g	n
Aluminium	Al	s	m
Manganese	Mn	s	m
Sodium	Na	s	m
Calcium	Ca	s	m
Chromium	Cr	s	m
Sulphur	S	s	n
Gold	Au	s	m
Copper	Cu	s	m

Structured questions (28%)

1. Give the symbol of the first 20 elements in the periodic table. [10]

H							He
Li	Be	B	C	N	O	F	Ne
Na	Mg	Al	Si	P	S	Cl	Ar
K	Ca						

2. (a) Which element(s) is / are in gas state at room temperature? [2]

U & T, Both of their bp and mp are lower than the room temperature.

- (b) Which element could be aluminium? [2]

P, It is the only conductor and has high bp and mp.

- (c) Which element is a metalloid? [2]

Q, It has intermediate conductivity and high mp and bp.

- (d) Which element(s) would melt if placed in hot water? [2]

R, Its mp is lower than hot water while its bp is higher than that.

- (e) Which element is a liquid at -50°C ? [2]

T, Its mp is lower than -50°C while its bp is higher than -50°C .

3. (a) Which element has 7 electrons in its outermost shell? Why? [2]

X, It is the only coloured gas.

- (b) Which element has the lowest melting point? Why? [2]

Y, It is the in Group VIII / a noble gas.

- (c) Which element has the largest number of protons? How do you know that? [2]

Z, It has the largest atomic number.

- (d) What would you expect to see if a piece of potassium is cut and left exposed to W? Explain your answer. [2]

Potassium will burn simultaneously since K is very reactive that it will react with W.

The End