

Practice with Net Ionic Equations

For each problem write:

- 1) balanced molecular equations – include the (aq), (s) symbols – deciding ahead of time if the salt is soluble will help in writing the ionic form correctly
 - 2) total ionic equation
 - 3) net ionic equation
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1. Calcium chloride solution is added to a solution of sodium carbonate.
2. Sodium phosphate solution is added to a solution of cobalt (II) chloride.
3. Copper (II) sulfate solution is added to a solution of sodium hydroxide.
4. Silver nitrate solution is added to a solution of strontium chloride.
5. Lead (II) nitrate solution is added to a solution of lead (IV) chloride.

Hints about the correct answers for the net ionic equation:

1. There is only one compound on the product side.
2. The sum of the coefficients on the reactant side is 5
3. There is one coefficient of 2 on the reactant side.
4. Each ion or compound has the same coefficient
5. The sum of the coefficients on the reactant side is 6