

# Science(Chemistry)

## Ionic Bonding II

Name: \_\_\_\_\_( ) Class: \_\_\_\_\_ Date: \_\_\_\_\_

1 i) Sodium and chlorine combine to form \_\_\_\_\_

Electronic Configuration

Noble gas structures

Na	2.8.1	loses <u>one</u> electron to form <u>Na<sup>+</sup></u>	2.8
Cl	2.8.7	gains _____ electron to form _____	2.8. ____

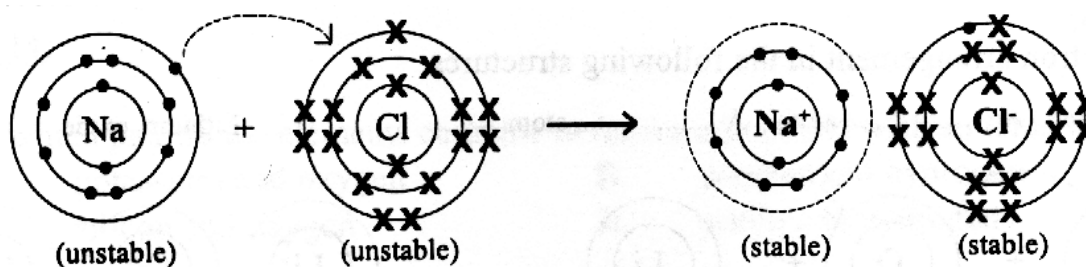
During combination, sodium ion (Na<sup>+</sup>) and \_\_\_\_\_ (Cl<sup>-</sup>) are formed.

ii) Electronic Structures:

A sodium atom

A chlorine atom

Sodium chloride



The **formula of sodium chloride** is ' Na<sup>+</sup>Cl<sup>-</sup> ' or ' NaCl '

2 i) Magnesium and oxygen combine to form \_\_\_\_\_

Electronic Configuration

Noble gas structures

Mg	_____	loses _____ electron to form _____	2.8
O	_____	gains _____ electron to form _____	2. ____

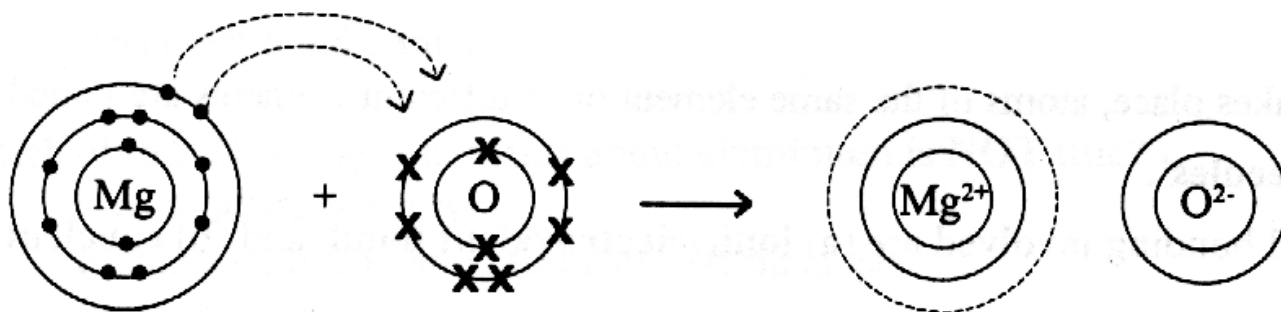
During combination, \_\_\_\_\_ ion (Mg<sup>2+</sup>) and \_\_\_\_\_ ion (O<sup>2-</sup>) are formed.

ii) Electronic Structures:

A \_\_\_\_\_ atom

An \_\_\_\_\_ atom

Magnesium Oxide



The **formula of magnesium oxide** is ' Mg<sup>2+</sup>O<sup>2-</sup> ' or ' \_\_\_\_\_ '

3 i) Calcium and Fluorine combine to form \_\_\_\_\_

Electronic Configuration

Noble gas structures

Ca \_\_\_\_\_ loses \_\_\_\_\_ electron to form \_\_\_\_\_ 2. \_\_\_\_\_

F \_\_\_\_\_ gains \_\_\_\_\_ electron to form \_\_\_\_\_ 2. \_\_\_\_\_

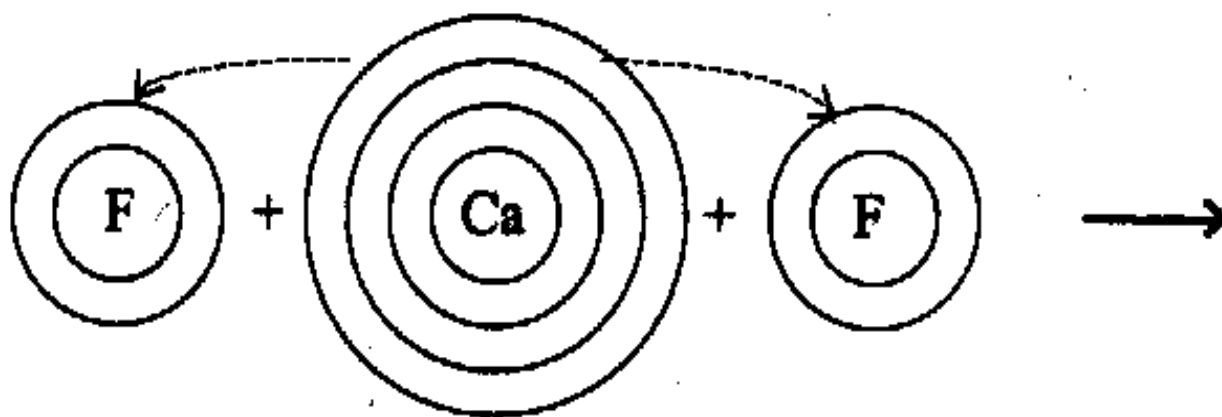
During combination, \_\_\_\_\_ ion ( $\text{Ca}^{2+}$ ) and \_\_\_\_\_ ion ( $\text{F}^-$ ) are formed.

cii) Electronic Structures:

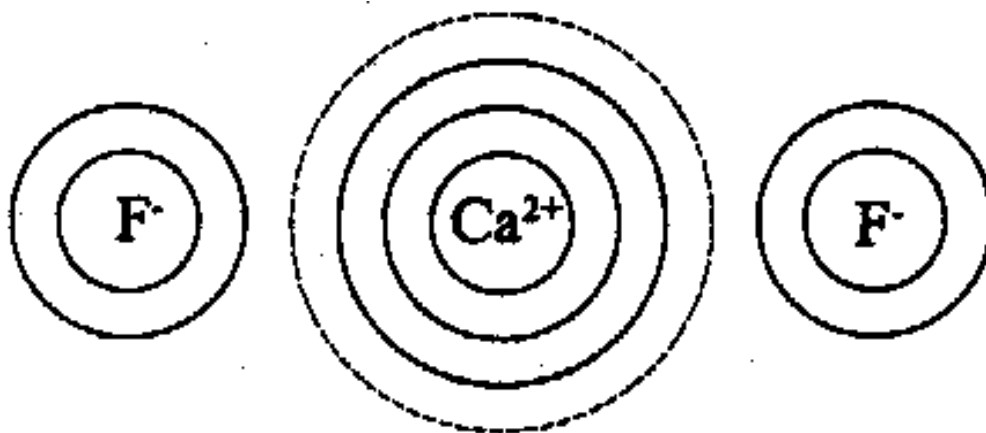
A \_\_\_\_\_ atom

A \_\_\_\_\_ atom

A fluorine atom



Calcium fluoride



The formula of calcium fluoride is ' \_\_\_\_\_ ' or ' \_\_\_\_\_ '

4 i) Lithium and Oxygen combine to form \_\_\_\_\_

Electronic Configuration

Noble gas structures

Li \_\_\_\_\_ loses \_\_\_\_\_ electron to form \_\_\_\_\_ 2. \_\_\_\_\_

O \_\_\_\_\_ gains \_\_\_\_\_ electron to form \_\_\_\_\_ 2. \_\_\_\_\_

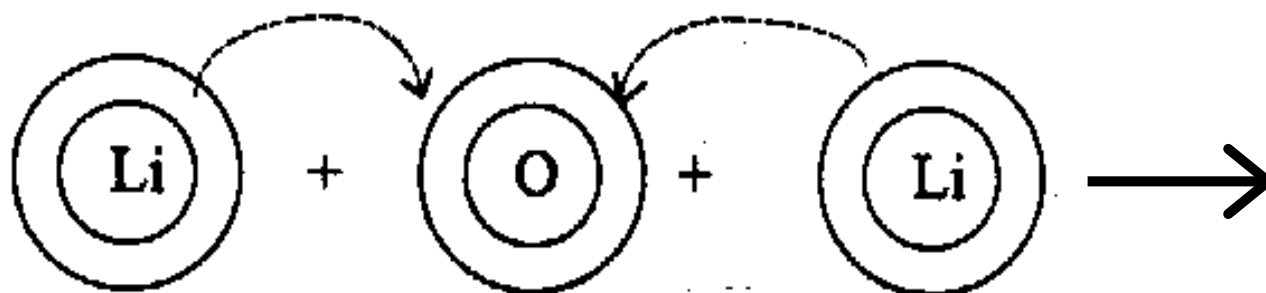
During combination, \_\_\_\_\_ ion ( $\text{Li}^+$ ) and \_\_\_\_\_ ion ( $\text{O}^{2-}$ ) are formed.

ii) Electronic Structures:

A \_\_\_\_\_ atom

A \_\_\_\_\_ atom

A \_\_\_\_\_ atom



Lithium oxide



The formula of lithium oxide is ' \_\_\_\_\_ ' or ' \_\_\_\_\_ '