

SCIENCE (Chemistry)

Acids and Bases - Exercise (Soluble and Insoluble Salts)

Name: _____ () Class: _____ Date: _____

1. Complete the following equations

- a) zinc + dilute sulphuric acid $\rightarrow$ _____ + _____
- b) potassium + dilute hydrochloric acid $\rightarrow$ _____ + _____
- c) calcium carbonate + dilute hydrochloric acid $\rightarrow$ _____ + _____ + _____
- d) copper(II) carbonate + dilute nitric acid $\rightarrow$ _____ + _____ + _____
- e) sodium carbonate + dilute sulphuric acid $\rightarrow$ _____ + _____ + _____
- f) copper(II) oxide + dilute nitric acid $\rightarrow$ _____ + _____
- g) calcium oxide + dilute sulphuric acid $\rightarrow$ _____ + _____
- h) sodium hydroxide + dilute hydrochloric acid $\rightarrow$ _____ + _____
- i) iron (III) hydroxide + dilute nitric acid $\rightarrow$ _____ + _____
- j) calcium hydroxide + dilute sulphuric acid $\rightarrow$ _____ + _____
- k) iron(II) oxide + dilute hydrochloric acid $\rightarrow$ _____ + _____
- l) iron(III) oxide + dilute sulphuric acid $\rightarrow$ _____ + _____
- m) copper(II) carbonate + dilute hydrochloric acid $\rightarrow$ _____ + _____ + _____

2. Fill in the blanks: Which of these salts are soluble? Which are insoluble?

- 1) Silver chlorides; 2) sodium chloride; 3) calcium carbonate; 4) ammonium sulphate;
5) potassium carbonate; 6) calcium nitrate; 7) Barium sulphate; 8) calcium chloride;
9) barium nitrates; 10) zinc carbonate; 11) magnesium chloride; 12) calcium sulphate;
13) zinc chloride; 14) lead(II) chlorides; 15) sodium carbonate; 16) calcium sulphate;
17) lead(II) sulphate; 18) magnesium carbonate; 19) ammonium carbonate;
20) copper(II) nitrate; 21) ammonium nitrate; 22) potassium nitrate; 23) iron(II) sulphate;
24) copper(II) sulphate; 25) zinc nitrate; 26) ammonium chloride; 27) silver nitrate
28) copper(II) carbonate; 29) copper(II) chloride; 30) iron(II) carbonate; 31) iron(II) chloride

Salts	Soluble	Insoluble
Chlorides		
Sulphates		
Carbonates		
Nitrates		
sodium salts		
Ammonium salts		
potassium salts		