

Name:

Atomic structure and isotopes

1. Write down the symbol for rubidium in the space below. Show the atomic number and the mass number. Next to the symbol write down how many protons, neutrons and electrons an atom of rubidium has.

2. There are 3 isotopes of the element with 1 proton. The most common one is called hydrogen. The other isotopes are called deuterium and tritium. Deuterium has one more neutron than hydrogen, and tritium has two more neutrons than hydrogen. Write down the symbols for hydrogen, deuterium and tritium, showing the mass numbers and atomic numbers.

3. Complete the table below

Atom	No. of protons	No. of neutrons	No. of electrons
${}^{79}_{35}\text{Br}$			
${}^{81}_{35}\text{Br}$			
${}^{14}_6\text{C}$			
${}^{56}_{26}\text{Fe}$			
${}^{235}_{92}\text{U}$			
${}^{238}_{92}\text{U}$			

